Chemistry 4th Nine Weeks: Scope and Sequence

Content Standards	Dates Taught	% of Students	Dates Re-taught (Optional)	Formative and Summative Assessments/ (Any Additional Comments Ontional)
		70% and	(optional)	optionaly
		over		
ACOS (4) Describe solubility in terms of energy changes associated				
with the solution process.				
Using solubility curves to interpret saturation levels				
Describing acids and bases in terms of strength, concentration, pH,				
and neutralization reactions				
Solving problems involving molarity, including solution preparation				
and dilution				
ALOS (5) Use the kinetic theory to explain states of matter, phase				
changes, solubility, <u>and chemical reactions</u> .				
Example:				
ACOS (6) Solve stoichiometric problems involving relationships among				
in a chamical reaction				
In a chemical reaction.				
reacting ionic and covalent bond types and products given known				
Assigning ovidation numbers for individual atoms of monatomic and				
nolvatomicions				
Identifying the nomendature of ionic compounds, binary				
compounds, and acids				
Classifying chemical reactions as composition, decomposition, single replacement, or double replacement				
Determining the empirical or molecular formula for a compound				
using percent composition [chapter 12]				
ACOS (7) Explain the behavior of ideal gases I terms of pressure,				
volume, temperature, and number of particles using Charles's law,				
Boyle's law, Gay –Lussac's law, the combined gas law, and the ideal				
qas law,				
ACOS (8) Distinguish among endothermic 9(g)-226.0(e)2.9(n)-2.R.0(S)-2	(I)]TJ 3-	2.9(e)2.9xr	[.] 9(r)-5.0r9(r)-	5.0r9(r)-signi8.0(r)-h.0(,)-222.0(si

Using LeChatlier's principle to explain changes in physical and chemical equilibrium		
ACOS (9) Distinguish between chemical and nuclear reactions. Identify atomic and subatomic particles Calculate half-life of selective radioactive isotopes Contrast fusion and fission Identify types of radiation and their properties		